Balance the following equations:

1) 
$$K_2O + H_2O \longrightarrow KOH$$

2) Al + FeO 
$$\longrightarrow$$
 Al<sub>2</sub>O<sub>3</sub> + Fe

3) NaOH + 
$$Cl_2$$
 NaCl +  $H_2O$ 

4) 
$$CO_2 + NH_3 \longrightarrow OC(NH_2)_2 + H_2O$$

5) 
$$N_2 + O_2 \longrightarrow N_2O$$

Complete the following reactions

1) 
$$P_4 + O_2 = -$$

3) Na + O<sub>2</sub> 
$$---$$

5) 
$$P_2O_5 + H_2O$$
 —

How will you prove that Aluminium oxide is amphoteric oxide?

Prove that sodium oxide is basic

Prove that oxides of sulphur is acidic

Write <u>THREE</u> separate equations to prepare oxygen in the laboratory.
1)
2)
3)
Draw a diagram of the apparatus for any one of above laboratory preparation method, you suggested

How oxygen is prepared in the industry. Give a brief account of it with a diagram of the process.